

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
12 July 2001 (12.07.2001)

PCT

(10) International Publication Number
WO 01/49695 A1

- (51) International Patent Classification⁷: **C07F 7/18** (74) Agent: **MA, Shirley, S.**; Crompton Corporation, 199 Benson Road, Middlebury, CT 06749 (US).
- (21) International Application Number: **PCT/US01/00242** (81) Designated States (*national*): BR, CN, JP, SG.
- (22) International Filing Date: **3 January 2001 (03.01.2001)** (84) Designated States (*regional*): European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, TR).
- (25) Filing Language: **English**
- (26) Publication Language: **English**
- (30) Priority Data:
09/478,283 **5 January 2000 (05.01.2000)** **US**
- (71) Applicant: **CROMPTON CORPORATION [US/US];**
199 Benson Road, Middlebury, CT 06749 (US).
- (72) Inventors: **REEDY, James, D.**; Route 2, Box 303B, Bel-
pre, OH 45714-9753 (US). **HARTMAN, Kenneth, M.**;
306 Dodd Street, Middlebourne, WV 26149 (US).

Published:

- *With international search report.*
— *With amended claims.*

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.



WO 01/49695 A1

(54) Title: **SULFUR-CONTAINING SILANE AS COUPLING AGENTS**

(57) Abstract: A method is provided for preparing sulfur-containing silane coupling agents by reacting an alkoxysilane acetal with a sulfurating agent. The sulfur-containing silane coupling agents so prepared include certain novel mercapto silanes, and are useful in providing rubber tires with improved properties.

BEST AVAILABLE COPY

SULFUR-CONTAINING SILANE AS COUPLING AGENTS

Background of the Invention

Sulfur-containing silane coupling agents are useful in providing rubber,
5 including automotive tires, with improved properties, generally by coupling inorganic
fillers or fibers with the rubber matrix in a fashion which leads to the improved
properties. The sulfur-containing silane coupling agents which have achieved
commercial success to date have been produced by disadvantageous processes which
involve the handling of large quantities of chlorine-containing by-products. Thus,
10 there is an ongoing need in the art to prepare sulfur-containing silane coupling agents
safely in high yields and efficiencies by processes which do not involve chlorine-
containing intermediates or by-products.

Current large-scale commercial production of such sulfur-containing silane
coupling agents is based on the raw material trichlorosilane. Trichlorosilane is
15 reacted with either allyl chloride or vinyltoluene to provide the respective
intermediates, 3-chloropropyltrichlorosilane or (trichlorosilylethyl)-toluene. The
former is reacted with alcohol to produce a 3-chloropropyl-trialkoxysilane, which is
reacted with sodium hydrosulfide or sodium tetrasulfide to produce the desired
products, plus four equivalents of chlorine. The latter is reacted with sulfur
20 monochloride, and the sulfurated trichlorosilane intermediate is reacted with alcohol
to produce the desired product, plus five equivalents of chlorine.

Both of the above embodiments suffer from additional disadvantages in that
yields from the reaction of trichlorosilane with allyl chloride are well below

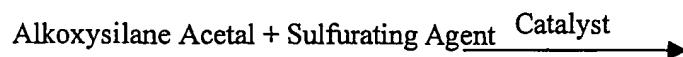
quantitative, based on the limiting reactant, with concurrent generation of undesired by-products. The reaction of trichlorosilane with vinyltoluene is susceptible to polymerization of the vinyltoluene, with subsequent reduced efficiency to (trichlorosilylethyl)toluene and formation of a polymeric by-product.

5 There is a need for a process that can simply and efficiently prepare certain known and novel silane coupling agents by a direct reaction, not producing chlorine-containing by-products. More specifically, there is a need for a process to produce sulfur-containing coupling agents that does not begin with trichlorosilane and does not involve chlorine-containing intermediates or by-products.

10

Summary of the Invention

The present invention provides a process for preparing sulfur-containing silane coupling agents, some of which are novel compositions, by the reactions of alkoxy silane acetals with nonionic, chlorine-free sulfurating agents including, but not
15 limited to, thiols, di- and higher thiols, hydrogen sulfide, and sulfur, in the presence of an acid catalyst, and optionally in the presence of hydrogen and an reduction catalyst. The process is depicted by the general reaction:



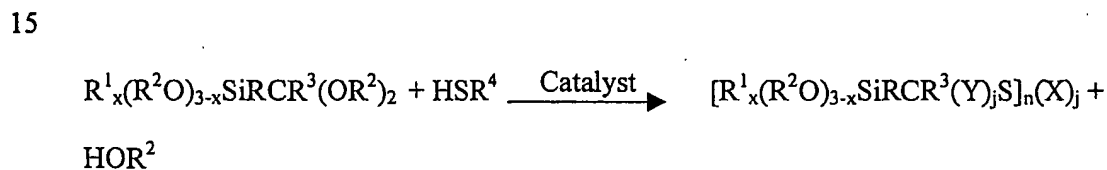
20

Sulfur-containing Silane Coupling Agent +

Detailed Description of the Invention

It is an object of the invention to provide an improved process for the synthesis of sulfur-containing silane compounds, useful as coupling agents, by reactions involving chlorine-free intermediates and sulfurating agents. It is another object of the invention to provide novel sulfur-containing silane compounds, useful as coupling agents. It is another object of the invention to provide a process for preparing intermediates for sulfur-containing silane compounds in high yields. It is yet another specific object of the invention to provide a process for preparing sulfur-containing silane compounds by a chlorine-free reaction that produces silanes without the production of undesired chlorine or polymeric intermediates or by-products.

It has been discovered that the acetal groups of alkoxysilane acetals can be reacted with sulfurating agents without causing significantly adverse reactions at the alkoxysilane groups. More particularly, the process of the invention is based on the following reaction:



where R is a saturated or unsaturated linear, branched, or cyclic divalent hydrocarbon group of 2 to 12 carbon atoms optionally containing divalent -O- or -S- linkages; R¹ is an alkyl group of 1 to 4 carbon atoms, an aryl group of 6 to 12 carbon atoms, or an aralkyl or alkaryl group of 7 to 13 carbon atoms; each R² can be R¹ or the 2 R² groups taken together form an R group so as to form a ring; R³ is a hydrogen atom or R¹; R⁴

is R^3 , $-RSH$, or $-RSiR^1_x(OR^2)_{3-x}$; x is an integer having a value of 0, 1, or 2, $j = 0$ or 1, n is an integer having a value of 1 to 4; when $n = 1$, $j = 1$, when $n > 1$, $j = 0$, X is R^4 and Y is $-SR^4$ or $-OR^2$, or alternatively X and Y taken together are $-SR-$ which forms a 1,3-dithiacycloalkane ring with the carbon bearing R^3 , when $n = 1$, or X and Y are both hydrogen atoms when $n = 1$ and R^4 is a hydrogen atom. The sulfur-containing silane coupling agents provided by the process of the present invention may thus be monomers or oligomers and may contain more than one molecular species. Where more than one group of R , R^1 , R^2 , R^3 , or R^4 is present in a molecule, said groups R , R^1 , R^2 , R^3 , and R^4 may be the same or different.

10 The alkoxysilane acetal raw materials are well-known in the art. The normal preparation is by hydrosilation of an unsaturated acetal, $CH_2=CR^3R^5CR^3(OR^2)_2$, with a hydroalkoxysilane, $R^1_x(R^2O)_{3-x}SiH$, in the presence of a catalyst,

$$CH_2=CR^3(R^5)_jCR^3(OR^2)_2 + R^1_x(R^2O)_{3-x}SiH \xrightarrow{\text{Catalyst}} R^1_x(R^2O)_{3-x}SiR^3CR^3(OR^2)_2$$

where j , R , R^1 , R^2 , and R^3 are as defined above; R^5 is a divalent saturated or

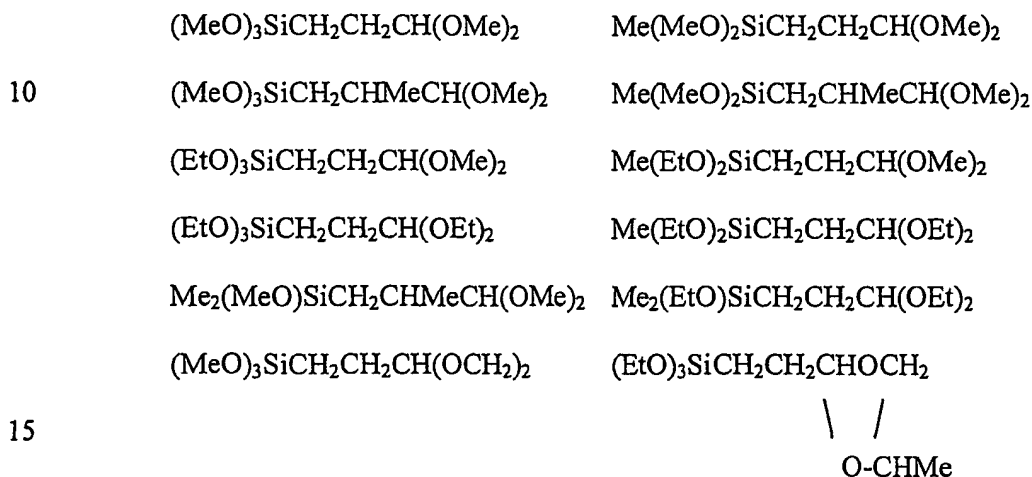
15 unsaturated linear, branched, or cyclic hydrocarbon group of 1 to 10 carbon atoms, optionally containing $-O-$ or $-S-$ linkages; the divalent group R being formed from the group $CH_2=CR^3R^5-$ upon hydrosilation.

The unsaturated acetals are articles of commerce, as are the hydroalkoxysilanes, and the catalysts used for hydrosilation. A particularly preferred

20 hydroalkoxysilane is trimethoxysilane, as prepared by the direct reaction between silicon metal and methanol.

Preferred alkoxysilane acetal raw materials include compounds wherein R is a linear or branched divalent hydrocarbon group of 2 to 4 carbon atoms, R^1 is an alkyl

group of 1 to 2 carbon atoms, R^2 is an alkyl group of 1 to 2 carbon atoms or 2 R^2 groups taken together form an R group, and R^3 is a hydrogen atom or an alkyl group of 1 to 2 carbon atoms. Most preferred versions of the alkoxysilane acetal include compounds wherein R is a linear divalent hydrocarbon group of 2 carbon atoms or a
 5 branched divalent hydrocarbon group of 3 carbon atoms, R^1 is a methyl group, R^2 is a methyl group or an ethyl group, and R^3 is a hydrogen or a methyl group. Thus, compounds in the most preferred group include the following:



Preferred sulfurating agents will depend on the product desired and may be selected singly or in combination from the group of thiols, di- and higher thiols, hydrogen sulfide, and sulfur, optionally in the presence of hydrogen and an effective
 20 reduction catalyst. Sulfur serves as a source of hydrogen sulfide in the presence of hydrogen and an effective reduction catalyst. Dithiols, including 1,2-dimercaptoethane, 1,2-dimercaptopropane, and 1,3-dimercaptopropane are particularly preferred for preparing certain sulfur-containing silane coupling agents which are novel compositions. The ratio between silane and sulfurating agent is not

narrowly critical, but preferably is between a slight excess of sulfurating agent to near stoichiometric equivalent.

A suitable reduction catalyst is selected from the group of metal-containing catalysts which are effective for reduction in the presence of sulfur and its compounds. Cobalt polysulfide is a preferred reduction catalyst for the process of the present invention, with a use level of 0.5 to 5 wt-% being preferred, and 1 to 3 wt-% being most preferred.

The process of the invention is preferably conducted in the presence of an acid catalyst, selected from the classes of protic (Bronsted) acids or nonprotic (Lewis) acids. The former is exemplified by para-toluenesulfonic acid and a wide variety of carboxylic and inorganic acids. The latter is exemplified by boron trifluoride, zinc salts, e.g., zinc chloride, and a number of other covalent metallic halides, including lanthanum chloride. Acids in solid or supported forms may also be used, including acid forms of zeolites, acid clays, sulfonated derivatives of fluoropolymers, and acids of either of the above classes deposited on inorganic supports. Selection of the acid catalyst is not a narrowly critical feature of the present invention, nor is its concentration. Any catalyst effective for the reaction can be used at an effective concentration. Acid concentrations in the range of 0.05 to 5 wt-% of the combined reactants are effective, with a concentration of 0.1 to 1 wt-% being preferred, and 0.1 to 0.5 wt-% being most preferred.

The process of the invention is performed at a temperature and pressure effective for the reaction, generally at an elevated temperature to assist in the removal of the alcohol by-product at ambient pressure. Preferred reaction temperatures are in

the range of 40° to 200°C and are not narrowly critical. The process may be run at atmospheric pressure for convenience, but may also be run at subatmospheric or superatmospheric pressures. Running at superatmospheric pressure will normally be necessary if a hydrogen atmosphere is maintained during the process in the presence of a reduction catalyst. Thus, equipment normally used for laboratory, pilot scale, or commercial scale synthetic chemistry, ranging from glassware to steel, may be used for the process of the present invention.

While the process of the present invention may be run in the absence of any solvent, solvents of various kinds may be used, for example, to assist in the introduction of raw materials and in the removal of the alcohol by-products. Among the available solvents are aromatic and aliphatic hydrocarbons, alcohols, ketones and ethers. Among the aromatic hydrocarbons are xylene, toluene, and benzene. Among the aliphatic hydrocarbons are pentane, hexane, heptane, octane, isooctane, decane, cyclohexane and methylcyclohexane. Among the alcohols are methanol, ethanol, isopropanol, propanol, butanol, hexanol, octanol and t-butanol. The ketones are represented by methyl ethyl ketone, methyl isopropyl ketone and cyclohexanone. The ethers are represented by tetrahydrofuran, dioxane, dioxolane and glyme. Certain of the solvents with low boiling points might require performing the reaction under elevated pressure.

Specific examples of products of the present invention may be selected, singly or in combinations, from the group of:

(MeO)₃Si(CH₂)₃SH; Me(MeO)₂Si(CH₂)₃SH; (EtO)₃Si(CH₂)₃SH;

Me(EtO)₂Si(CH₂)₃SH; Me₂(MeO)Si(CH₂)₃SH; (MeO)₃SiCH₂CHMeCH₂SH; and

Me(EtO)₂SiCH₂CHMeCH₂SH;

which are known compositions of matter, and from the group of novel mecapto silane

5 structures of:

$[R^1_x(R^2O)_{3-x}SiRC(R^3)_2S]_cR'Si(OR^2)_{3-x}R^1_x$ wherein R' is a c+1 valent hydrocarbon and
c = 1 to 3, preferably 2 and preferably at least one R³ is R¹;

$[R^1_x(R^2O)_{3-x}SiRC(R^3)(OR^2)S]_2R$;

$R^1_x(R^2O)_{3-x}SiRC(R^3)_2SRSH$;

10 $[R^1_x(R^2O)_{3-x}SiRCR^3(-)S(-)]_q$ wherein q ≥ 2, preferably 3; and

$R^1_x(R^2O)_{3-x}SiR-cyclo(C(R^3)SRS-)$;

Wherein R¹, R², R³, R⁴, R, and x are as above.

Examples of the foregoing novel structures are as follows:

15

(MeO)₃Si(CH₂)₂CH(SCH₂)₂; (EtO)₃Si(CH₂)₂CHSCH₂CHMeS;

(MeO)₃Si(CH₂)₂CH(SCH₂)₂CH₂

Me(MeO)₂Si(CH₂)₂CH(SCH₂)₂; (MeO)₃SiCH₂CHMeCH(SCH₂)₂;

(MeO)₃Si(CH₂)₃SCH=CHCH₂Si(OMe)₃; [(MeO)₃Si(CH₂)₃S]₂CH(CH₂)₂Si(OMe)₃;

20 (MeO)₃Si(CH₂)₂CH(OMe)S(CH₂)₂SH; [(MeO)₃Si(CH₂)₂CHS]₃;

[Me(MeO)₂Si(CH₂)₂CHS]₃; [(MeO)₃Si(CH₂)₂CH(OMe)SCH₂]₂;

[(EtO)₃Si(CH₂)₂CHS]₃; and (MeO)₃Si(CH₂)₃S(CH₂)₂SH.

The reaction mixtures produced by the process of this invention will typically comprise a mixture of subject silanes and alcohol. Depending on the particular alcohol, transesterification with the alkoxy silane may occur. The alcohol may be removed by distillation or vacuum stripping. It is preferred to separate the alcohol to
5 reduce its concentration to less than about 0.5%. There is typically no need to separate the individual silanes; however, where desired, this can be accomplished by chromatography or high or ultra-high-vacuum distillation.

The products prepared according to the invention are employed in natural and synthetic rubber compositions and blends of known and novel formulation, in
10 amounts consistent with those previously employed for other silane coupling agents for the use in sulfur-vulcanizable, silica-reinforced tire rubber compositions. Exemplary of suitable amounts will be at least 2 parts per hundred parts rubber (PHR) and, preferably from about 4 to about 20 PHR, *e.g.*, 6 to 12 PHR. The amount will also be related to the amount of silica employed, preferably the ratio by weight of
15 silica to silane being in the range of from 4:1 to about 40:1, more narrowly from about 6:1 to about 10:1. Molar ratios of added sulfur for vulcanization to sulfur in the silane can be varied within the range of from above 0 to about 100:1 or more, preferably from 2:1 to 20:1, more narrowly from 5:1 to 10:1. The required amount of silane will decrease as its relative sulfur content increases.

20 Exemplary of suitable rubber compositions are sulfur-vulcanizable synthetic rubber compositions. Representative examples of suitable rubber polymers include solution styrene-butadiene rubber (SSBR), styrene-butadiene rubber (SBR), natural rubber (NR), polybutadiene (BR), ethylene-propylene co- and ter- polymers (EP,

EPDM), and acrylonitrile-butadiene rubber (NBR). The rubber composition preferably is comprised of at least one diene-based elastomer, or rubber. Suitable conjugated dienes are isoprene and 1,3-butadiene and suitable vinyl aromatic compounds are styrene and alpha methyl styrene. Thus, the rubber is a sulfur curable rubber. Such diene based elastomer, or rubber, may be selected, for example, from at least one of cis-1,4-polyisoprene rubber (natural and/or synthetic), and preferably natural rubber), emulsion polymerization prepared styrene/butadiene copolymer rubber, organic solution polymerization prepared styrene/butadiene rubber, 3,4-polyisoprene rubber, isoprene/butadiene rubber, styrene/isoprene/butadiene terpolymer rubber, cis-1,4-polybutadiene, medium vinyl polybutadiene rubber (35-50 percent vinyl), high vinyl polybutadiene rubber (50-75 percent vinyl), styrene/isoprene copolymers, emulsion polymerization prepared styrene/butadiene/acrylonitrile terpolymer rubber and butadiene/acrylonitrile copolymer rubber. An emulsion polymerization derived styrene/butadiene (E-SBR) might be used having a relatively conventional styrene content of 20 to 28 percent bound styrene or, for some applications, an E-SBR having a medium to relatively high bound styrene content, namely, a bound styrene content of 30 to 45 percent. Emulsion polymerization prepared styrene/butadiene/acrylonitrile terpolymer rubbers containing 2 to 40 weight percent bound acrylonitrile in the terpolymer are also contemplated as diene based rubbers for use in this invention.

See any of U. S. Patent Nos. 3,451,458; 5,110,969, 5,227,425 and 5,753,732, for examples of rubber compounds that can be improved with the invention with silica

as a reinforcing agent. The disclosures of these patents are incorporated by reference in their entireties.

The rubber compositions, in addition to at least one elastomer of synthetic or natural origin, will contain a mineral filler, particularly silica, in amounts effective for reinforcing the rubber in its vulcanized state. The silica can be of the types known, for example described in U.S. Patent Nos. 4,704,414, 5,227,425 and 5,753,732, and will be employed in amounts suitable for reinforcing tires, especially those having low rolling resistance. The silica will be employed at a level of from about 5 to about 100 parts per hundred parts of rubber, preferably at least 30 parts silica. Higher or lesser amounts can be employed where appropriate.

Precipitated silicas are preferred fillers. The silica may be characterized by having a BET surface area, as measured using nitrogen gas, preferably in the range of 40 to 600, and more usually in a range of 50 to 300 m²/g. The silica typically may also be characterized by having a dibutylphthalate (DBP) absorption value in a range of 100 to 350, and more usually 150 to 300. Further, the silica, as well as the aforesaid alumina and aluminosilicate, may be expected to have a CTAB surface area in a range of 100 to 220. The average mercury porosity specific surface area for the silica should be in a range of 100 to 300 m²/g.

The rubber composition may be compounded by methods known in the rubber compounding art, such as mixing the various sulfur-vulcanizable constituent rubbers with various commonly used additive materials such as, for example, curing aids, such as sulfur, activators, retarders and accelerators, processing additives, such as oils, resins including tackifying resins, silicas, plasticizers, fillers, pigments, fatty acid,

zinc oxide, waxes, antioxidants and antiozonants, peptizing agents, and reinforcing materials such as, for example, carbon black. Depending on the intended use of the sulfur vulcanizable and sulfur vulcanized material (rubbers), the additives mentioned above are selected and commonly used in conventional amounts.

5 A rubber composition may prepared by a process such as by:

(A) thermomechanically mixing, in at least one preparatory mixing step, to a temperature of 140 °C to 200 °C, alternatively to 140 °C to 190 °C, for a total mixing time of 2 to 20, alternatively 4 to 15, minutes for such mixing step(s)

(i) 100 parts by weight of at least one sulfur vulcanizable rubber selected from
10 conjugated diene homopolymers and copolymers, and copolymers of at least one conjugated diene and aromatic vinyl compound, (ii) 5 to 100, preferably 25 to 80, phr (parts per hundred rubber) of particulate filler, wherein preferably the filler contains 1 to 85 weight percent carbon black (iii) 0.05 to 20 parts by weight filler of at least one sulfur-containing silane;

15 (B) subsequently blending therewith, in a final thermomechanical mixing step at a temperature to 50 °C to 130 °C for a time sufficient to blend the rubber, preferably between 1 to 30 minutes, more preferably 1 to 3 minutes, a curing agent at 0 to 5 phr; and optionally

(C) curing said mixture at a temperature of 130 to 200 °C for about 5 to 60
20 minutes.

An exemplary process for using silane coupling agents to manufacture silica containing rubber is disclosed in PCT/US98/17391, which is incorporated herein by reference.

The rubber compositions of the invention are employed to form various rubber articles, including shoe soles and tire parts, such as treads and sidewalls in the normal fashion as conventional silica-reinforced, sulfur vulcanizable rubber compositions.

Whereas the exact scope of this invention is set forth in the appended claims,
5 the following specific examples illustrate certain aspects of the present invention and, more particularly, point out various aspects of the method for evaluating same.

However, the examples are set forth for illustrative purposes only and are not to be construed as limitations on the present invention. The abbreviations g, ml, mm, m, psi, ppm, GC, MS, and NMR respectively represent gram, milliliter, millimeter, molar
10 equivalent, pounds per square inch, parts per million, gas chromatography, mass spectrometry, and nuclear magnetic resonance spectroscopy; temperature is reported in degrees Centigrade. Unless stated otherwise, all reactions were run in standard laboratory glassware of various sizes at atmospheric pressure under an inert atmosphere of nitrogen, and all parts and percentages are by weight.

15 Alkoxysilane acetals were prepared by the platinum-catalyzed hydrosilations of acrolein dimethyl acetal, acrolein 1,2-propylene acetal, and methacrolein dimethyl acetal with trimethoxysilane or methyldimethoxysilane.

20 **EXAMPLE 1 - PREPARATION OF 2-(2-TRIMETHOXYSYLYLETHYL)-1,3-DITHIOLANE**

To a 250 ml three-necked flask fitted with a mechanical stirrer and a distillation head were added 73.6 g of 95.2% (3,3-dimethoxypropyl)trimethoxysilane (70.07 g, 0.313 m), 32.4 g (0.344 m) of ethane-1,2-dithiol, and 0.2 g of anhydrous

zinc chloride catalyst. The reaction was heated to 130° under nitrogen and maintained at that temperature for 1.75 hours, during which time 18.7 g of 95.4% methanol (89% of theory) were distilled from the flask. The crude residue, containing 90% of the title product, was distilled to provide 55.7 g of 94.3% pure product (70.1% yield). The product was characterized by GC/MS and NMR. Similar results were obtained when 0.2 wt-% of zinc hexacyanocobaltate was used as catalyst in place of zinc chloride.

EXAMPLE 2 - REACTION OF (3,3-DIMETHOXYPROPYL)TRIMETHOXY-SILANE WITH ETHANE-1,2-DITHIOL USING LANTHANUM CHLORIDE

The procedure and apparatus of Example 1 were used with 77.2 g (0.345 m) of (3,3-dimethoxypropyl)trimethoxysilane, 37.5 g (0.371 m) of ethane-1,2-dithiol and 0.23 g of lanthanum chloride as catalyst. After distillation of methanol, the crude product (89.6 g) contained 20% of methanol and raw materials, plus 14.5% of 2-(2-trimethoxysilylethyl)-1,3-dithiolane, 37.3% of the monosubstitution product, $(\text{MeO})_3\text{Si}(\text{CH}_2)_2\text{CH}(\text{OMe})\text{S}(\text{CH}_2)_2\text{SH}$, and 19.6% of a dimeric product, $[(\text{MeO})_3\text{Si}(\text{CH}_2)_2\text{CH}(\text{OMe})\text{SCH}_2]_2$, as determined by GC/MS.

EXAMPLE 3 - PREPARATION OF 2-(2-METHYLDIMETHOXY-SILYLETHYL)-1,3-DITHIOLANE

The procedure and apparatus of Example 1 were used with 36.2 g (0.177 m) of (3,3-dimethoxypropyl)methyldimethoxysilane, 18.31 g (0.195 m) of ethane-1,2-dithiol, and 0.13 g of zinc chloride catalyst. Reaction and work-up as in Example 1

provided crude product containing lesser amounts of raw materials and 69.4% of the title compound, as determined by GC/MS.

EXAMPLE 4 - PREPARATION OF 2-(2-TRIMETHOXYSILYLETHYL)-1,3-

5 DITHIANE

The procedure and apparatus of Example 1 were used with 22.4 g (0.207 m) of propane-1,3-dithiol, 46.4 g (0.207 m) of (3,3-dimethoxypropyl)trimethoxysilane, 0.14 g of p-toluene-sulfonic acid, and 100 ml of xylenes solvent. Methanol was removed by distillation as the contents of the flask were heated at reflux (143°) for 2 hours.

- 10 The crude product, excluding solvent, contained 75% of the title product and 7.4% of the monosubstituted product, $(\text{MeO})_3\text{Si}(\text{CH}_2)_2\text{CH}(\text{OMe})\text{S}(\text{CH}_2)_3\text{SH}$, as determined by GC/MS.

EXAMPLE 5 - PREPARATION OF 2-(1-METHYL-2-

15 TRIMETHOXYSILYLETHYL)-1,3-DITHIOLANE

The procedure and apparatus of Example 1 were used with 25.0 g (0.102 m) of (3,3-dimethoxy-2-methylpropyl)trimethoxysilane (boiling point, 125°/25 mm, prepared by the hydrosilation of methacrolein dimethyl acetal with trimethoxysilane, as characterized by GC/MS and NMR), 9.6 g (0.102 m) of ethane-1,2-dithiol, and

20 0.07 g of anhydrous zinc chloride catalyst. The crude product, after distillation to a pot temperature of 195°, was 27.1 g, containing 64.3% of the title product as determined by GC/MS.

EXAMPLE 6 - HYDROGENOLYSIS OF A DITHIOLANE

Into a 300 ml rocking autoclave were charged 58.2 g (0.331 m) of 2-(n-pentyl)-1,3-dithiolane, 1.06 g of sulfur, 0.12 g of p-toluenesulfonic acid, and 1.87 g of cobalt polysulfide catalyst (added as a 43% slurry in toluene). The autoclave was pressurized with hydrogen and maintained between 1600 and 2200 psi while being heated to 250°. After cooling and venting, the crude product contained 57% of the hydrogenolysis product, $\text{CH}_3(\text{CH}_2)_5\text{S}(\text{CH}_2)_2\text{SH}$, as determined by GC/MS. When this procedure is applied to the product of Example 1, the hydrogenolysis product, $(\text{MeO})_3\text{Si}(\text{CH}_2)_3\text{S}(\text{CH}_2)_2\text{SH}$, is obtained.

EXAMPLE 7 - REACTION OF ALKOXYSILANE ACETAL WITH H₂S

An apparatus was assembled consisting of a 500 ml three-necked flask fitted with a mechanical stirrer, distillation head and receiver, sparge tube to introduce hydrogen sulfide, and a sodium methylate scrubber to remove hydrogen sulfide from the exiting gases. To the flask were charged 108.3 g (0.426 m) of (3,3-dimethoxypropyl)-trimethoxysilane, 0.3 g of p-toluenesulfonic acid, and 200 g of toluene solvent. The flask and contents were heated at 100-120° for 9 hours while hydrogen sulfide was introduced at 0.04 liters/minute. The flask contents were heated to 140° and volatiles (including 3-mercaptopropyltrimethoxysilane) removed at 1 mm vacuum. The crude product contained several components, including 4.3% of

$[(\text{MeO})_3\text{SiCH}_2\text{CH}=\text{CH}]_2\text{S}$ and 30.5% of cis/trans-

$(\text{MeO})_3\text{Si}(\text{CH}_2)_2\text{CH}(\text{OMe})\text{SCH}=\text{CHCH}_2\text{Si}(\text{OMe})_3$ as determined by GC/MS.

EXAMPLE 8 - REACTION OF 3-MERCAPTOPROPYLTRIMETHOXYSILANE

5 WITH

(3,3-DIMETHOXYPROPYL)TRIMETHOXYSILANE

The procedure and apparatus of Example 1 were used with 98.1 g (0.50 m) of 3-
.mercaptopropyltrimethoxysilane, 52.5 g (0.227 m) of (3,3-dimethoxypropyl)-
trimethoxysilane, and 0.75 g of zinc chloride catalyst. The reactants were heated at
10 135°C for 2.75 hours and 11.8 g (93.3% methanol) of volatiles were removed. The
crude product, 129 g, was analyzed by GC/MS and contained 27.3% 3-
mercaptopropyl-trimethoxysilane, 35% cis/trans- $(\text{MeO})_3\text{Si}(\text{CH}_2)_3\text{S}-$
 $\text{CH}=\text{CHCH}_2\text{Si}(\text{OMe})_3$, and 12% of the dithioacetal
 $[(\text{MeO})_3\text{Si}(\text{CH}_2)_3\text{S}]_2\text{CH}(\text{CH}_2)_2\text{Si}(\text{OMe})_3$, as determined by GC/MS.

15

EXAMPLE 9 - PREPARATION OF 3-

MERCAPTOPROPYLTRIMETHOXYSILANE AND A TRITHIANE FROM AN ALKOXYLSILANE ACETAL AND SULFUR

A series of experiments was conducted using a 300 ml rocking autoclave,
20 charged with varying amounts of (3,3-dimethoxypropyl)trimethoxysilane (42.6-195
g), sulfur (1.05-1.2 m), cobalt or nickel polysulfide catalyst (2.5 wt-%), acetic acid or
p-toluenesulfonic acid (0.5 wt-% or 0.2 wt-%, respectively), and hydrogen (1600 psi),

with heat applied at 150-250° for 4-23 hours. Triethylamine and tributylamine at 0.5 wt-% were also screened in place of the acid component. Good results were obtained, for example, when run with 42.6 g of acetal, 1.05 m of sulfur, acetic acid, and cobalt polysulfide, wherein the yield of 3-mercaptopropyltrimethoxysilane, as determined by GC/MS of the crude product, was 29%. The major product, of several higher boiling components, has been assigned the trithiane structure, $[(\text{MeO})_3\text{Si}(\text{CH}_2)_2\text{CHS}]_3$, based on spectral comparisons to a more volatile model compound.

EXAMPLE 10 - PREPARATION OF 3-

MERCAPTOPROPYLTRIMETHOXYSILANE AND A TRITHIANE FROM AN ALKOXYSILANE ACETAL AND HYDROGEN SULFIDE

Certain of the experiments of Example 9 were repeated using hydrogen sulfide rather than sulfur as the source of sulfur. Good results were obtained when run with 121.5 g of the alkoxy silane acetal, 0.5 wt-% of acetic acid, 2.5 wt-% of cobalt polysulfide catalyst, 34g (1.85 m) of hydrogen sulfide, and hydrogen at 1600 psi as charged at room temperature. Heating up to 200° and holding for 2.25 hours provided 158.3 g of crude product from which a 61% yield of 3-mercaptopropyltrimethoxysilane (based on acetal) was recovered by distillation. Nondistilled product was largely the same trithiane as observed in Example 9.

The above description is intended to enable the person skilled in the art to practice the invention. It is not intended to detail all of the possible modifications and variations which will become apparent to the skilled worker upon reading the description. It is

intended, however, that all such modifications and variations be included within the scope of the invention which is defined by the following claims. The claims are meant to cover the indicated elements and steps in any arrangement or sequence which is effective to meet the objectives intended for the invention, unless the context

5 specifically indicates the contrary.

CLAIMS

- 1 1. A method for preparing a sulfur silane comprising reacting an alkoxysilane acetal
2 with a sulfurating agent in the presence of a catalyst.

- 1 2. The method of Claim 1 wherein said sulfur-containing silane has the formula
2 $[R^1_x(R^2O)_{3-x}SiR^3CR^3(Y)_jS]_n(X)_j$ where R is a saturated or unsaturated linear,
3 branched, or cyclic divalent hydrocarbon group of 2 to 12 carbon atoms optionally
4 containing divalent -O- or -S- linkages; R^1 is an alkyl group of 1 to 4 carbon
5 atoms, an aryl group of 6 to 12 carbon atoms, or an aralkyl or alkaryl group of 7 to
6 13 carbon atoms; each R^2 can be R^1 or the 2 R^2 groups taken together form an R
7 group so as to form a ring; R^3 is a hydrogen atom or R^1 ; R^4 is R^3 , -RSH, or -
8 $RSiR^1_x(OR^2)_{3-x}$; x is an integer having a value of 0, 1, or 2, n is an integer having a
9 value of 1 to 4; j = 0 or 1, when n = 1, j = 1, when n > 1, j = 0, X is R^4 and Y is -
10 SR^4 or $-OR^2$, or alternatively X and Y taken together are -SR- which forms a
11 1,3-dithiacycloalkane ring with the carbon bearing R^3 when n = 1, or X and Y are
12 both hydrogen atoms when n = 1 and R^4 is a hydrogen atom.

- 1 3. The method of Claim 2 wherein R is a saturated linear or branched divalent
2 hydrocarbon group of 2 or 3 carbon atoms; R^1 is a methyl group; R^2 is R^1 , an ethyl
3 group, or 2 R^2 groups taken together form an R group; R^3 is a hydrogen atom or a
4 methyl group; R^4 is a hydrogen atom, the group -RSH, or the group -
5 $RSiR^1_x(OR^2)_{3-x}$; and x is an integer having a value of 0 or 1; n is an integer having
6 a value of 1 or 3.

- 1 4. The method of Claim 2 wherein said alkoxysilane acetal is selected from the
2 group consisting of $(\text{MeO})_3\text{SiCH}_2\text{CH}_2\text{CH}(\text{OMe})_2$,
3 $\text{Me}(\text{MeO})_2\text{SiCH}_2\text{CH}_2\text{CH}(\text{OMe})_2$, $(\text{MeO})_3\text{SiCH}_2\text{CHMeCH}(\text{OMe})_2$,
4 $\text{Me}(\text{MeO})_2\text{SiCH}_2\text{CHMeCH}(\text{OMe})_2$, $(\text{EtO})_3\text{SiCH}_2\text{CH}_2\text{CH}(\text{OMe})_2$,
5 $\text{Me}(\text{EtO})_2\text{SiCH}_2\text{CH}_2\text{CH}(\text{OMe})_2$, $(\text{EtO})_3\text{SiCH}_2\text{CH}_2\text{CH}(\text{OEt})_2$, and
6 $\text{Me}(\text{EtO})_2\text{SiCH}_2\text{CH}_2\text{CH}(\text{OEt})_2$, and said sulfurating agent is selected from the
7 group of hydrogen sulfide, HSRSH and $\text{HSRSiR}^1_x(\text{OR}^2)_{3-x}$ and sulfur.

- 1 5. The method of Claim 1 wherein the catalyst is selected from the group of
2 Bronsted acids and Lewis acids in liquid or solid form and is at a concentration of
3 0.05 to 5 wt-% of the combined reactants.

- 1 6. The method of Claim 1 wherein the sulfurating agent is hydrogen sulfide or sulfur
2 in the presence of hydrogen gas and a reduction catalyst and the major product is
3 $\text{HSRSiR}^1_x(\text{OR}^2)_{3-x}$ where R is a saturated linear or branched divalent hydrocarbon
4 group, R^1 is a methyl group, R^2 is a methyl group or an ethyl group, and x is an
5 integer having a value of 0 or 1.

- 1 7. The method of Claim 6 wherein the reduction catalyst is cobalt polysulfide and
2 the major product is $\text{HS}(\text{CH}_2)_3\text{Si}(\text{OMe})_3$.

- 1 8. The method according to Claim 7 additionally comprising the step of removing
2 alcohol from the sulfur silane following the reaction.

- 1 9. A silane selected from the group consisting of:

- 2 $[R^1_x(R^2O)_{3-x}SiRC(R^3)_2S]_cR'Si(OR^2)_{3-x}R^1_x$ wherein R' is a $c+1$ valent hydrocarbon and
 3 $c = 1$ to 3 ;
 4 $[R^1_x(R^2O)_{3-x}SiRC(R^3)(OR^2)S]_2R$;
 5 $R^1_x(R^2O)_{3-x}SiRC(R^3)_2SRSH$;
 6 $[R^1_x(R^2O)_{3-x}SiRCR^3(-)S(-)]_q$ wherein $q \Rightarrow 2$, and
 7 $R^1_x(R^2O)_{3-x}SiR-cyclo(C(R^3)SRS-)$.

1 10. A silane according to claim 9 of the formula $[R^1_x(R^2O)_{3-x}SiRC(R^3)_2S]_cR'Si(OR^2)_{3-x}R^1_x$ wherein R' is a $c+1$ valent hydrocarbon and $c = 1$ to 3 .

1 11. A silane according to claim 9 of the formula $[R^1_x(R^2O)_{3-x}SiRC(R^3)(OR^2)S]_2R$.

1 12. A silane according to claim 9 of the formula $R^1_x(R^2O)_{3-x}SiRC(R^3)_2SRSH$.

1 13. A silane according to claim 9 of the formula $[R^1_x(R^2O)_{3-x}SiRCR^3(-)S(-)]_q$ wherein
 2 $q \Rightarrow 2$.

1 14. A silane according to claim 9 of the formula $R^1_x(R^2O)_{3-x}SiR-cyclo(C(R^3)SRS-)$.

1 15. A silane according to claim 9 selected from the group consisting of

- 2 $(MeO)_3Si(CH_2)_2CH(SCH_2)_2$; $(EtO)_3Si(CH_2)_2CHSCH_2CHMeS$;
 3 $(MeO)_3Si(CH_2)_2CH(SCH_2)_2CH_2$; $Me(MeO)_2Si(CH_2)_2CH(SCH_2)_2$;
 4 $(MeO)_3SiCH_2CHMeCH(SCH_2)_2$; $(MeO)_3Si(CH_2)_3SCH=CHCH_2Si(OMe)_3$;
 5 $[(MeO)_3Si(CH_2)_3S]_2CH(CH_2)_2Si(OMe)_3$; $(MeO)_3Si(CH_2)_2CH(OMe)S(CH_2)_2SH$;
 6 $[(MeO)_3Si(CH_2)_2CHS]_3$; $[Me(MeO)_2Si(CH_2)_2CHS]_3$;

7 $[(\text{MeO})_3\text{Si}(\text{CH}_2)_2\text{CH}(\text{OMe})\text{SCH}_2]_2$; $[(\text{EtO})_3\text{Si}(\text{CH}_2)_2\text{CHS}]_3$; and

8 $(\text{MeO})_3\text{Si}(\text{CH}_2)_3\text{S}(\text{CH}_2)_2\text{SH}$.

1 16. A rubber composition comprising rubber, silica and a silane of claim 9.

AMENDED CLAIMS

[received by the International Bureau on 8 May 2001 (08.05.01);
original claims 1-16 replaced by new claims 1-14 (4 pages)]

1 1. A method for preparing a sulfur silane comprising reacting an
2 alkoxysilane acetal with a sulfurating agent in the presence of a catalyst, wherein
3 said sulfur-containing silane has the formula $[R^1_x(R^2O)_{3-x}SiR^3CR^4(Y)_j]_n(X)_j$
4 where
5 R is a saturated or unsaturated linear, branched, or cyclic divalent
6 hydrocarbon group of 2 to 12 carbon atoms optionally containing divalent -O- or
7 -S- linkages;
8 R¹ is an alkyl group of 1 to 4 carbon atoms, an aryl group of 6 to 12
9 carbon atoms, or an aralkyl or alkaryl group of 7 to 13 carbon atoms;
10 each R² can be R¹ or the 2 R² groups taken together form an R group so
11 as to form a ring;
12 R³ is a hydrogen atom or R¹;
13 R⁴ is R³, -RSH, or -RSiR¹_x(OR²)_{3-x};
14 x is an integer having a value of 0, 1, or 2;
15 n is an integer having a value of 1 to 4;
16 j = 0 or 1, when n = 1, j = 1, when n > 1, j = 0, X is R⁴ and Y is -SR⁴
17 or -OR², or alternatively X and Y taken together are -SR- which forms a
18 1,3-dithiacycloalkane ring with the carbon bearing R³ when n = 1, or X and Y
19 are both hydrogen atoms when n = 1 and R⁴ is a hydrogen atom.

1 2. The method of claim 1 wherein R is a saturated linear or
2 branched divalent hydrocarbon group of 2 or 3 carbon atoms; R¹ is a methyl
3 group; R² is R¹, an ethyl group, or 2 R² groups taken together form an R group;
4 R³ is a hydrogen atom or a methyl group; R⁴ is a hydrogen atom, the group
5 -RSH, or the group -RSiR¹_x(OR²)_{3-x}; and x is an integer having a value of 0 or 1;
6 n is an integer having a value of 1 or 3.

1 3. The method of claim 1 wherein said alkoxysilane acetal is selected
2 from the group consisting of (MeO)₃SiCH₂CH₂CH(OMe)₂,
3 Me(MeO)₂SiCH₂CH₂CH(OMe)₂, (MeO)₃SiCH₂CHMeCH(OMe)₂,
4 Me(MeO)₂SiCH₂CHMeCH(OMe)₂, (EtO)₃SiCH₂CH₂CH(OMe)₂,

5 Me(EtO)₂SiCH₂CH₂CH(OMe)₂, (EtO)₃SiCH₂CH₂CH(OEt)₂, and
 6 Me(EtO)₂SiCH₂CH₂CH(OEt)₂, and said sulfurating agent is selected from the
 7 group of hydrogen sulfide, HSRSH and HRSiR¹_x(OR²)_{3-x} and sulfur.

1 4. The method of claim 1 wherein the catalyst is selected from the
 2 group of Bronsted acids and Lewis acids in liquid or solid form and is at a
 3 concentration of 0.05 to 5 weight percent of the combined reactants.

1 5. The method of claim 1 wherein the sulfurating agent is hydrogen
 2 sulfide or sulfur in the presence of hydrogen gas and a reduction catalyst and
 3 the major product is HRSiR¹_x(OR²)_{3-x} where R is a saturated linear or
 4 branched divalent hydrocarbon group, R¹ is a methyl group, R² is a methyl
 5 group or an ethyl group, and x is an integer having a value of 0 or 1.

1 6. The method of claim 5 wherein the reduction catalyst is cobalt
 2 polysulfide and the major product is HS(CH₂)₃Si(OMe)₃.

1 7. The method according to claim 6 additionally comprising the step
 2 of removing alcohol from the sulfur silane following the reaction.

1 8. A silane made according to the method of claim 1, said silane
 2 being a member selected from the group consisting of:
 3 [R¹_x(R²O)_{3-x}SiRC(R³)₂S]_cR'Si(OR²)_{3-x}R¹_x wherein R' is a c+1 valent hydrocarbon
 4 and c = 1 to 3;
 5 [R¹_x(R²O)_{3-x}SiRC(R³)(OR²)S]₂R;
 6 R¹_x(R²O)_{3-x}SiRC(R³)₂SRSH;
 7 [R¹_x(R²O)_{3-x}SiRCR³(-)S(-)]_q wherein q = >2; and
 8 R¹_x(R²O)_{3-x}SiR-cyclo(C(R³)SRS-).

1 9. A silane consisting of a formula $R^1_x(R^2O)_{3-x}SiRC(R^3)_2SRSH$ where
 2 R is a saturated or unsaturated linear, branched, or cyclic divalent
 3 hydrocarbon group of 2 to 12 carbon atoms optionally containing divalent -O- or
 4 -S- linkages;
 5 R¹ is an alkyl group of 1 to 4 carbon atoms, an aryl group of 6 to 12
 6 carbon atoms, or an aralkyl or alkaryl group of 7 to 13 carbon atoms;
 7 each R² can be R¹ or the 2 R² groups taken together form an R group so
 8 as to form a ring;
 9 R³ is a hydrogen atom or R¹;
 10 R⁴ is R³, -RSH, or $-RSiR^1_x(OR^2)_{3-x}$; and
 11 x is an integer having a value of 0, 1, or 2.

1 10. A silane consisting of the formula
 2 $R^1_x(R^2O)_{3-x}SiR-cyclo(C(R^3)SRS-)$ where
 3 R is a saturated or unsaturated linear, branched, or cyclic divalent
 4 hydrocarbon group of 2 to 12 carbon atoms optionally containing divalent -O- or
 5 -S- linkages;
 6 R¹ is an alkyl group of 1 to 4 carbon atoms, an aryl group of 6 to 12
 7 carbon atoms, or an aralkyl or alkaryl group of 7 to 13 carbon atoms;
 8 each R² can be R¹ or the 2 R² groups taken together form an R group so
 9 as to form a ring;
 10 R³ is a hydrogen atom or R¹;
 11 R⁴ is R³, -RSH, or $-RSiR^1_x(OR^2)_{3-x}$; and
 12 x is an integer having a value of 0, 1, or 2.

1 11. A silane consisting of a formula selected from the group consisting
 2 of $(MeO)_3Si(CH_2)_2CH(SCH_2)_2$; $(EtO)_3Si(CH_2)_2CHSCH_2CHMeS$;
 3 $(MeO)_3Si(CH_2)_2CH(SCH_2)_2CH_2$; $Me(MeO)_2Si(CH_2)_2CH(SCH_2)_2$;
 4 $(MeO)_3SiCH_2CHMeCH(SCH_2)_2$; $(MeO)_3Si(CH_2)_3SCH=CHCH_2Si(OMe)_3$;
 5 $[MeO]_3Si(CH_2)_3S]_2CH(CH_2)_2Si(OMe)_3$; $(MeO)_3Si(CH_2)_2CH(OMe)S(CH_2)_2SH$;
 6 $[(MeO)_3Si(CH_2)_2CHS]_3$; $[Me(MeO)_2Si(CH_2)_2CHS]_3$;

- 7 [(MeO)₃Si(CH₂)₂CH(OMe)SCH₂]₂; [(EtO)₃Si(CH₂)₂CHS]₃; and
8 (MeO)₃Si(CH₂)₃S(CH₂)₂SH.

1 12. A rubber composition comprising rubber, silica, and a silane of
2 claim 9.

1 13. A rubber composition comprising rubber, silica, and a silane of
2 claim 10.

1 14. A rubber composition comprising rubber, silica, and a silane of
2 claim 11.

INTERNATIONAL SEARCH REPORT

Intern Application No

PCT/US 01/00242

A. CLASSIFICATION OF SUBJECT MATTER
IPC 7 C07F/18

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 C07F

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

WPI Data, PAJ, EPO-Internal, CHEM ABS Data

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	PERARD, SERGE ET AL: "18-Functionalized steroids: synthesis of thio derivatives of progesterone" STERIODS, vol. 55, no. 6, 1990, pages 271-275, XP000982546 page 273	1
X	EP 0 507 727 A (CIBA GEIGY AG) 7 October 1992 (1992-10-07)	9, 11
A	page 10; example 12	1
X	EP 0 963 995 A (SHINETSU CHEMICAL CO) 15 December 1999 (1999-12-15)	9, 10, 13, 16
A	page 5	1
	--- -/--	

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

* Special categories of cited documents:

- *A* document defining the general state of the art which is not considered to be of particular relevance
- *E* earlier document but published on or after the international filing date
- *L* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- *O* document referring to an oral disclosure, use, exhibition or other means
- *P* document published prior to the international filing date but later than the priority date claimed

- *T* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- *X* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- *Y* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- *G* document member of the same patent family

Date of the actual completion of the international search

27 February 2001

Date of mailing of the international search report

08/03/2001

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2
NL - 2280 HV Rijswijk
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,
Fax: (+31-70) 340-3016

Authorized officer

Bader, K

INTERNATIONAL SEARCH REPORT

Inter 1al Application No

PCT/US 01/00242

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 5 286 815 A (LEIR CHARLES M ET AL) 15 February 1994 (1994-02-15)	9,11
A	column 15; examples 8,9 ----	1
X	WO 99 15583 A (VON HELLENS CARL WALTER ;BAYER INC (CA)) 1 April 1999 (1999-04-01)	9,10,16
A	page 19 -page 20 ----	1
X	PATENT ABSTRACTS OF JAPAN vol. 1998, no. 04, 31 March 1998 (1998-03-31) & JP 09 316330 A (POLA CHEM IND INC), 9 December 1997 (1997-12-09) abstract ----	9,11
X	PATENT ABSTRACTS OF JAPAN vol. 1997, no. 11, 28 November 1997 (1997-11-28) & JP 09 176323 A (POLA CHEM IND INC), 8 July 1997 (1997-07-08) abstract ----	9,11
X	BUDER, WOLFGANG: "Synthesis and spectroscopic study of trimethoxysilylpropylsulfanes" Z. NATURFORSCH., B: ANORG. CHEM., ORG. CHEM., vol. 34B, no. 6, 1979, pages 790-793, XP000986985 ----	9,10,13
A	page 792 -page 793 ----	1
P,X	EP 1 035 162 A (SHINETSU CHEMICAL CO) 13 September 2000 (2000-09-13) page 5 -page 6 ----	9,10,16
A	US 5 916 973 A (MATERNE THIERRY FLORENT EDME ET AL) 29 June 1999 (1999-06-29) column 11; claim 9 -----	9,16

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/US 01/00242

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
EP 0507727	A	07-10-1992	US 5266715 A	30-11-1993
			AU 658392 B	13-04-1995
			AU 1139392 A	10-09-1992
			BR 9200744 A	10-11-1992
			CA 2062312 A	07-09-1992
			DE 69216893 D	06-03-1997
			DE 69216893 T	25-09-1997
			ES 2099239 T	16-05-1997
			JP 5125084 A	21-05-1993
			KR 199343 B	15-06-1999
			MX 9200911 A	01-09-1992
			US 5372889 A	13-12-1994
			US 5426204 A	20-06-1995
EP 0963995	A	15-12-1999	JP 11349594 A	21-12-1999
			US 6114560 A	05-09-2000
US 5286815	A	15-02-1994	AU 3479893 A	03-09-1993
			BR 9305848 A	18-02-1997
			CA 2128865 A	19-08-1993
			DE 69306378 D	16-01-1997
			DE 69306378 T	26-06-1997
			EP 0625176 A	23-11-1994
			JP 7503748 T	20-04-1995
			MX 9300556 A	01-09-1993
			WO 9316142 A	19-08-1993
WO 9915583	A	01-04-1999	AU 8795798 A	12-04-1999
			BR 9812659 A	22-08-2000
			EP 1021484 A	26-07-2000
			PL 339261 A	04-12-2000
			ZA 9808559 A	30-03-1999
JP 09316330	A	09-12-1997	NONE	
JP 09176323	A	08-07-1997	NONE	
EP 1035162	A	13-09-2000	JP 2000265062 A	26-09-2000
US 5916973	A	29-06-1999	BR 9701229 A	10-11-1998
			CA 2197832 A	07-09-1997
			EP 0794188 A	10-09-1997
			JP 10007687 A	13-01-1998